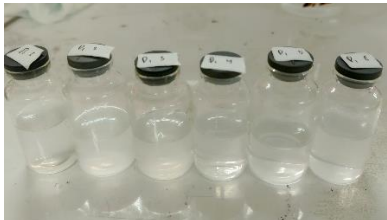
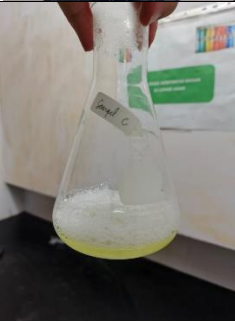


LAMPIRAN

Lampiran 1. Pembuatan larutan

Keterangan	Gambar
Penimbangan baku Natrium Sulfat Anhidrat	
Pembuatan kurva baku 6 seri konsentrasi	
Validasi metode akurasi	
Validasi metode presisi	
Penetapan kadar Sulfat	

Lampiran 2. Preparasi sampel

Keterangan	Gambar
Penimbangan sampel	
Sampel dilarutkan dengan 50 mL aquades dengan suhu 70°C dan ditambahkan H ₂ O ₂ 30% 2 mL	
Sampel setelah penambahan 1 mL HCL 1% dan sampel setelah didiamkan	
Disaring kemudian ditambahkan aquades sampai tanda batas	

Lampiran 3. Perhitungan kurva kalibrasi

Baku sulfat menggunakan natrium sulfat anhidrat (Na_2SO_4), perhitungan yang ditimbang:

$$\begin{aligned}
 &= \frac{\text{BM Na}_2\text{SO}_4}{\text{BM SO}_4} = \frac{\text{g Na}_2\text{SO}_4}{\text{g SO}_4} \\
 &= \frac{142 \text{ g/mol}}{96 \text{ g/mol}} = \frac{\text{g Na}_2\text{SO}_4}{0,1 \text{ g}} \\
 &= \frac{142 \text{ g/mol}}{96 \text{ g/mol}} = \frac{\text{g Na}_2\text{SO}_4}{0,1 \text{ g}} = 0,1479 \text{ gr Na}_2\text{SO}_4
 \end{aligned}$$

Baku Na_2SO_4 147,9 mg/100 mL $\rightarrow$ 1000 $\mu\text{g/mL}$ SO_4

Perhitungan larutan standar SO_4 100 $\mu\text{g/mL}$

$$V_1 \times C_1 = V_2 \times C_2$$

$$V_1 \times 1000 = 50 \text{ mL} \times 100$$

$$V_1 = \frac{50 \times 100}{1000}$$

$$V_1 = 5 \text{ mL}$$

Perhitungan kurva baku 6 seri konsentrasi:

$$20 \mu\text{g/mL}: V_1 \times C_1 = V_2 \times C_2$$

$$V_1 \times 100 = 10 \text{ mL} \times 20$$

$$V_1 = \frac{20 \times 10}{100}$$

$$V_1 = 2 \text{ mL}$$

$$30 \mu\text{g/mL}: V_1 \times C_1 = V_2 \times C_2$$

$$V_1 \times 100 = 10 \text{ mL} \times 30$$

$$V_1 = \frac{30 \times 10}{100}$$

$$V_1 = 3 \text{ mL}$$

$$40 \mu\text{g/mL}: V_1 \times C_1 = V_2 \times C_2$$

$$V_1 \times 100 = 10 \text{ mL} \times 40$$

$$V_1 = \frac{40 \times 10}{100}$$

$$V_1 = 4 \text{ mL}$$

$$50 \text{ }\mu\text{g/mL: } V1 \times C1 = V2 \times C2$$

$$V1 \times 100 = 10 \text{ mL} \times 50$$

$$V1 = \frac{50 \times 10}{100}$$

$$V1 = 5 \text{ mL}$$

$$60 \text{ }\mu\text{g/mL: } V1 \times C1 = V2 \times C2$$

$$V1 \times 100 = 10 \text{ mL} \times 60$$

$$V1 = \frac{60 \times 10}{100}$$

$$V1 = 6 \text{ mL}$$

$$70 \text{ }\mu\text{g/mL: } V1 \times C1 = V2 \times C2$$

$$V1 \times 100 = 10 \text{ mL} \times 70$$

$$V1 = \frac{70 \times 10}{100}$$

$$V1 = 7 \text{ mL}$$

Lampiran 4. Perhitungan linieritas dan sensitivitas

Persamaan linieritas; $y = bx + a$

$$y = 0,0097x + 0,0113$$

$$a = 0,0113$$

$$b = 0,0097$$

$$r = 0,9993$$

No.	Konsentrasi	Abs	Y'	(Y-Y') ²
1.	20	0,204	0,2053	0,00000169
2.	30	0,311	0,3023	0,00007569
3.	40	0,394	0,3993	0,00002809
4.	50	0,498	0,4963	0,00000289
5.	60	0,587	0,5933	0,00003969
6.	70	0,698	0,6903	0,00005929
$\bar{x} = 45$				$\Sigma = 0,00020734$

$$\frac{Sy}{x} = \sqrt{\frac{\Sigma(y-y')^2}{n-2}} = \sqrt{\frac{0,00020734}{4}} = 0,007199653$$

$$BD = \frac{3 \times (Sy/x)}{b} = \frac{3 \times 0,007199653}{0,0097} = 2,226696733 \text{ } \mu\text{g/mL}$$

$$BK = \frac{10 \times (Sy/x)}{b} = \frac{10 \times 0,007199653}{0,0097} = 7,422322443 \text{ } \mu\text{g/mL}$$

$$Sx0 = \frac{Sy/x}{b} = \frac{0,007199653}{0,0097} = 0,742232244$$

$$Vx0 = \frac{Sx0}{\bar{x}} \times 100 = \frac{0,742232244}{45} \times 100 = 1,649404987\%$$

Lampiran 5. Perhitungan akurasi

No	Abs adisi	C adisi	Abs sampel	C sampel	C adisi – C sampel	C teoritis	Rata-rata ± SD
1	0,479	48,21649	0,311	30,89690	17,76632	20	89,51 ± 2,64
2	0,486	48,93814	0,301	29,86597	18,48797	20	
3	0,476	47,90722	0,308	30,58762	17,45704	20	
			Rata-rata	30,45017			
1	0,578	58,42268	0,311	30,89690	27,97250	30	91,40 ± 2,09
2	0,574	58,01031	0,301	29,86597	27,56013	30	
3	0,566	57,18557	0,308	30,58762	26,73539	30	
			Rata-rata	30,45017			
1	0,672	68,11340	0,311	30,8969	37,66323	40	93,38 ± 0,77
2	0,666	67,49485	0,301	29,8659	37,04467	40	
3	0,669	67,80412	0,308	30,5876	37,35395	40	
			Rata-rata	30,45017			

Rumus perhitungan:

Menghitung konsentrasi (x) = $\frac{Y-a}{b}$

% Recovery = $\frac{\text{Konsentrasi adisi} - \text{Konsentrasi sampel}}{\text{C teoritis}} \times 100\%$

Lampiran 6. Perhitungan presisi

Hari	Abs adisi	C Adisi	C sampel	C adisi – C sampel	Rata-rata	SD	% RSD
Ke-1	0,565	57,082474	30,793814	26,28865979	26,76976	0,47309	1,76726
	0,568	57,391753		26,59793814			
	0,567	57,288665		26,49484536			
	0,578	58,422680		27,62886598			
	0,574	58,010309		27,21649485			
	0,566	57,185567		26,39175258			
Ke-2	0,571	57,701031	30,793814	26,90721649	26,76976	0,47309	1,76726
	0,568	57,391753		26,59793814			
	0,562	56,773196		25,97938144			
	0,575	58,113402		27,31958763			
	0,571	57,701031		26,90721649			
	0,566	57,185567		26,39175258			
Ke-3	0,570	57,597938	30,793814	26,80412371	26,76976	0,47309	1,76726
	0,572	57,804124		27,01030928			
	0,568	57,391753		26,59793814			
	0,563	56,876289		26,08247423			
	0,577	58,319588		27,5257732			
	0,573	57,907216		27,11340206			

Lampiran 7. Penetapan kadar

Sampel A

Pengulangan	Abs	C ($\mu\text{g/mL}$)	FP	FK	V pelarut (mL)	W sampel (mg)	Kadar (μg)	Kadar (mg)	%Kadar
1	0,516	52,031	2	1,09	100	100	3793,548	3,793548	3,793548
2	0,509	51,309	2	1,09	100	100	3740,933	3,740933	3,740933
3	0,511	51,515	2	1,09	100	100	3755,966	3,755966	3,755966
								Rata-rata	3,763
								SD	0,027

Sampel B

Pengulangan	Abs	C ($\mu\text{g/mL}$)	FP	FK	V pelarut (mL)	W sampel (mg)	Kadar (μg)	Kadar (mg)	%Kadar
1	0,321	31,928	2	1,09	100	100	2327,842	2,327842	2,327842
2	0,325	32,340	2	1,09	100	100	2357,908	2,357908	2,357908
3	0,316	31,412	2	1,09	100	100	2290,26	2,29026	2,29026
								Rata-rata	2,325
								SD	0,033

Rumus perhitungan kadar sulfur:

$$\%S = \frac{\frac{BMS}{BMSO_4} \times \text{Masa (mg)}}{\text{Sampel yang ditimbang}} \times 100\%$$

Lampiran 8. Bukti bimbingan

NIM	23IFF04030	Nama Mahasiswa	AMMAR SURI DHARMAWAN
Program Studi	SI Farmasi	Jenis TA	Skripsi
Periode Mulai	2024 Genap	SKS Lulus	150 SKS
Tgl. Mulai	4 Maret 2025	Judul Tugas Akhir	Analisis Kandungan Sulfur pada Sediaan Sabun Menggunakan Metode Turbidimetri
Tahap	Sidang Akhir/ Komprehensif (Ujian)	Status	Aktif

No	Tanggal	Dosen Pembimbing	Topik	Disetujui	Aksi
1	21 April 2025	apt. Anne Yuliantini, M.Si.	Diskusi metode analisis	✓	 
1	20 Maret 2025	Dr. apt. Winasih Rachmawati, M.Si.	Diskusi Metode yang digunakan	✓	 
2	2 Mei 2025	apt. Anne Yuliantini, M.Si.	Kemajuan 1	✓	 
2	17 April 2025	Dr. apt. Winasih Rachmawati, M.Si.	Konsul Hasil Penelitian	✓	 
3	13 Juni 2025	apt. Anne Yuliantini, M.Si.	Bimbingan dan Kemajuan 2	✓	 
3	2 Mei 2025	Dr. apt. Winasih Rachmawati, M.Si.	Kemajuan 1	✓	 
4	23 Juni 2025	apt. Anne Yuliantini, M.Si.	Melakukan bimbingan Hasil	✓	 
4	6 Mei 2025	Dr. apt. Winasih Rachmawati, M.Si.	Konsultasi hasil Linearitas dan Sensitifitas	✓	 
5	30 Juni 2025	apt. Anne Yuliantini, M.Si.	Melakukan bimbingan Draft BAB 4 dan BAB 5	✓	 
5	2 Juni 2025	Dr. apt. Winasih Rachmawati, M.Si.	Konsultasi hasil dari Absorbansi sampel	✓	 
6	2 Juli 2025	apt. Anne Yuliantini, M.Si.	Draft Skripsi	✓	 
7	4 Juli 2025	apt. Anne Yuliantini, M.Si.	Bimbingan draft	✓	 
7	16 Juli 2025	Dr. apt. Winasih Rachmawati, M.Si.	Bimbingan Prosedur Kerja akurasi, presisi	✓	 
8	18 Juni 2025	Dr. apt. Winasih Rachmawati, M.Si.	Hasil olah data	✓	 
8	9 Juli 2025	apt. Anne Yuliantini, M.Si.	Revisi draft skripsi	✓	 
9	24 Juli 2025	Dr. apt. Winasih Rachmawati, M.Si.	revisi Draft Skripsi	✓	 
10	4 Juli 2025	Dr. apt. Winasih Rachmawati, M.Si.	Revisi draft skripsi	✓	 

Lampiran 9. Hasil turnitin

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Lampiran 10. *Curriculum vitae*

CURRICULUM VITAE	
PERSONAL DATA Nama : AMMAR SURI DHARMAWAN Tempat, Tgl. Lahir : Bandar Lampung, 3 Mei 2001 Jenis Kelamin : Laki-laki Tinggi Badan : 173 cm Berat Badan : 63 kg Alamat Perum. Bukit Kemiling Permai No 130 Kec Kemiling Permai, Kota Bandar Lampung Telepon : 085278014095 Email : amarsuri79@gmail.com Sosial Media : IG @amrsuri.d	
RIWAYAT PENDIDIKAN 2005 TK Al-Azhar 16 Bandar Lampung 2006 – 2012 SD Al-Kautsar Bandar Lampung 2013 – 2016 SMP Negeri 22 Bandar Lampung 2016 – 2019 SMA Negeri 14 Bandar Lampung 2019 – 2021 D3 Farmasi Politeknik Kesehatan Tanjungkarang 2023 – 2025 S1 Farmasi Universitas Bhakti Kencana	
PENGALAMAN ORGANISASI 2019 – 2021 Badan Eksekutif Mahasiswa Poltekkes Tanjungkarang 2019 – 2021 Korps Sukarela Palang Merah Indonesia tingkat Universitas	
KEMAMPUAN - Microsoft Word, Canva	