

LAMPIRAN**Lampiran 1. Tahapan Penelitian**

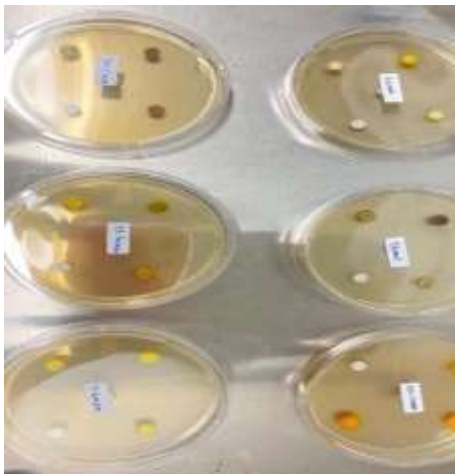
	
Pengumpulan Bahan	
	
Proses Pengeringan	
	
Serbuk Simplisia	

 A laboratory setup for reflux extraction. Two round-bottom flasks containing a yellowish liquid are placed on magnetic stirrers. The flasks are connected to a vertical glass tube that leads to a condenser, which is also connected to a collection flask. The setup is supported by metal stands and clamps.	Ekstraksi Menggunakan Refluks Rimpang dan Daun Kunyit
 A laboratory setup for evaporation. A round-bottom flask containing a dark liquid is placed on a magnetic stirrer. The flask is connected to a vertical glass tube that leads to a condenser, which is also connected to a collection flask. The setup is supported by metal stands and clamps.	Pemekatan Menggunakan Evaporasi
 A laboratory setup showing several small glass vials containing a yellowish liquid, arranged in a row. In the background, there are larger glass bottles and a blue tube.	Kuersetin
 A laboratory setup showing several small glass vials containing a yellowish liquid, arranged in a row. In the background, there are larger glass bottles and a blue tube.	Asam Galat

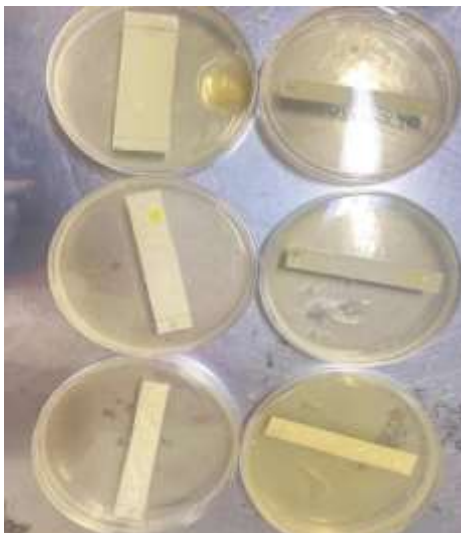
	Sterilisasi menggunakan Autoclave
	Nutrient Agar (NA)
	Peremajaan Bakteri <i>Propionibacterium acne</i>
	Inokulasi



Spektrofotometer Uv-Vis



Uji Aktivitas Antibakteri



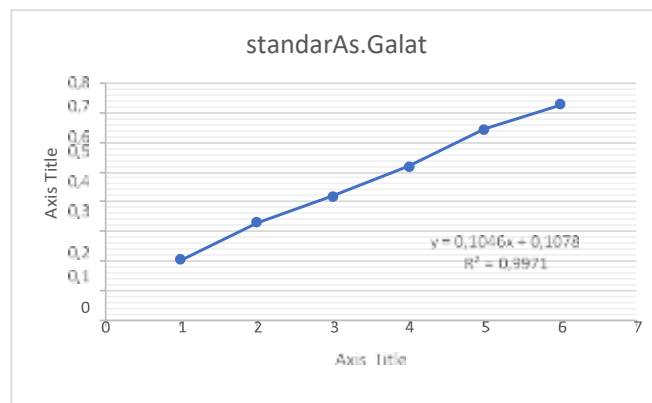
Uji Bioautografi

Lampiran 2. Pengujian Kadar Fenol

Kadar Fenol

Tabel 2.1 Hasil Kurva Kalibrasi Standar Asam Galat λ 765 nm

Konsentrasi (ppm)	Absorbansi
20 ppm	0,204
30 ppm	0,33
40 ppm	0,418
50 ppm	0,52
60 ppm	0,645
70 ppm	0,727



Hasil Pengamatan kadar fenol

Nilai a = 0,1078

Nilai b = 0,1046

Nilai $r^2 = 0,9971$

- Konsentrasi N-heksan Rimpang = 400 $\mu\text{g/mL}$, sampel 16mg/6,66mL

1. 0,4258

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4258-0,1453}{0,1033} = \frac{0,3180}{0,1033} = 3,040 \text{ mg/mL}$$

2. 0,4283

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4283-0,1453}{0,1033} = \frac{0,3205}{0,1033} = 3,064 \text{ mg/mL}$$

3. 0,4368

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4368-0,1453}{0,1033} = \frac{0,3290}{0,1033} = 3,145 \text{ mg/mL}$$

- % Kadar N-Heksan Rimpang

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{3,040 \times 10 \times 6,66}{16} = 12,655\%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{3,064 \times 10 \times 6,66}{16} = 12,754\%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{3,145 \times 10 \times 6,66}{16} = 13,092\%$$

- Konsentrasi Etil Asetat Rimpang = 600 $\mu\text{g/mL}$, sampel 24mg/8,62mL

1. 0,5541

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5541-0,1453}{0,1033} = \frac{0,4463}{0,1033} = 4,267 \text{ mg/mL}$$

2. 0,5619

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5619-0,1453}{0,1033} = \frac{0,4541}{0,1033} = 4,341 \text{ mg/mL}$$

3. 0,5736

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5736-0,1453}{0,1033} = \frac{0,4658}{0,1033} = 4,453 \text{ mg/mL}$$

- % Kadar Konsentrasi Etil Asetat Rimpang

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,267 \times 10 \times 8,62}{24} = 15,325\%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,341 \times 10 \times 8,62}{24} = 15,593\%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,453 \times 10 \times 8,62}{24} = 15,994\%$$

- Konsentrasi Etanol Rimpang = 400 µg/mL, sampel 16mg/5,71mL

1. 0,5533

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5533-0,1453}{0,1033} = \frac{0,4455}{0,1033} = 4,259 \text{ mg/mL}$$

2. 0,5396

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5396-0,1453}{0,1033} = \frac{0,4318}{0,1033} = 4,128 \text{ mg/mL}$$

3. 0,5341

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5341-0,1453}{0,1033} = \frac{0,4263}{0,1033} = 4,076 \text{ mg/mL}$$

- % Kadar Etanol Rimpang

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,259 \times 10 \times 5,71}{16} = 15,200\%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,128 \times 10 \times 5,71}{16} = 14,732\%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,076 \times 10 \times 5,71}{16} = 14,545\%$$

- Konsentrasi N-heksan Daun = 400 µg/mL, sampel 16mg/8mL

1. 0,3662

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3662-0,1453}{0,1033} = \frac{0,2584}{0,1033} = 2,470 \text{ mg/mL}$$

2. 0,3625

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3625-0,1453}{0,1033} = \frac{0,2547}{0,1033} = 2,435 \text{ mg/mL}$$

3. 0,3615

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3615-0,1453}{0,1033} = \frac{0,2537}{0,1033} = 2,425 \text{ mg/mL}$$

- % Kadar N-heksan Daun

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,470 \times 10 \times 8}{16} = 12,352\%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,435 \times 10 \times 8}{16} = 12,175\%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,425 \times 10 \times 8}{16} = 12,127\%$$

- Konsentrasi Etil Asetat Daun = 200 µg/mL, sampel 8mg/5mL

1. 0,2082

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2082-0,1453}{0,1033} = \frac{0,1004}{0,1033} = 0,960 \text{ mg/mL}$$

2. 0,2073

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2073-0,1453}{0,1033} = \frac{0,0995}{0,1033} = 0,951 \text{ mg/mL}$$

3. 0,2064

$$4. \text{ Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2064-0,1453}{0,1033} = \frac{0,0986}{0,1033} = 0,943 \text{ mg/mL}$$

- % Kadar Etil Asetat Daun

$$1. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{0,960 \times 10 \times 5}{8} = 5,999\%$$

$$2. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{0,951 \times 10 \times 5}{8} = 5,945\%$$

$$3. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{0,943 \times 10 \times 5}{8} = 5,891\%$$

- Konsentrasi Etanol Daun = 600 µg/mL sampel 24mg/10mL

1. 0,4343

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4343-0,1453}{0,1033} = \frac{0,3265}{0,1033} = 3,121 \text{ mg/mL}$$

2. 0,4235

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4235-0,1453}{0,1033} = \frac{0,3157}{0,1033} = 3,018 \text{ mg/mL}$$

3. 0,4162

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4162-0,1453}{0,1033} = \frac{0,3084}{0,1033} = 2,948 \text{ mg/mL}$$

- % Kadar Etanol Daun

$$1. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{3,121 \times 10 \times 10}{24} = 13,006\%$$

$$2. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{3,018 \times 10 \times 10}{24} = 12,576\%$$

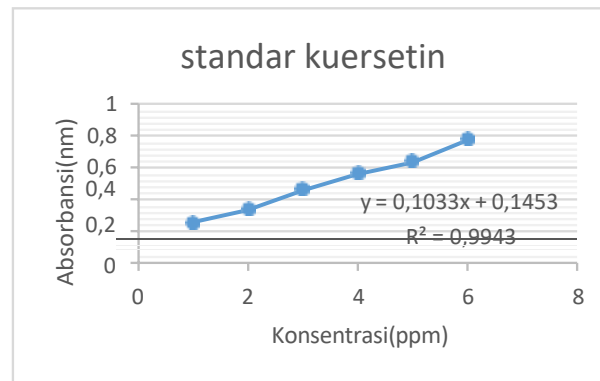
$$3. \text{ \% Kadar} = \frac{C \times V \times FP}{M} = \frac{2,948 \times 10 \times 10}{24} = 12,285\%$$

Lampiran 3. Pengujian Kadar Flavonoid

Kadar Flavonoid

Tabel 3.1 Hasil Kurva Kalibrasi Standar Kuersetin λ 420 nm

Konsentrasi (ppm)	Absorbansi
20 ppm	0,256
30 ppm	0,337
40 ppm	0,467
50 ppm	0,562
60 ppm	0,640
70 ppm	0,778



Hasil pengamatan kadar flavonoid

Nilai a = 0,1453

Nilai b = 0,1033

Nilai r^2 = 0,9943

- Konsentrasi N-heksan Rimpang = 400 $\mu\text{g/mL}$, sampel = 16mg/8mL

1. 0,3319

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3319-0,1453}{0,1033} = \frac{0,1866}{0,1033} = 1,806 \text{ mg/mL}$$

2. 0,3317

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3317-0,1453}{0,1033} = \frac{0,1864}{0,1033} = 1,804 \text{ mg/mL}$$

3. 0,3687

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,3687-0,1453}{0,1033} = \frac{0,2234}{0,1033} = 2,163 \text{ mg/mL}$$

- % Kadar N-heksan Rimpang

$$1. \% \text{ kadar} = \frac{C \times V \times FP}{M} = \frac{1,806 \times 10 \times 8}{16} = 9,032\%$$

$$2. \% \text{ kadar} = \frac{C \times V \times FP}{M} = \frac{1,804 \times 10 \times 8}{16} = 9,022\%$$

$$3. \% \text{ kadar} = \frac{C \times V \times FP}{M} = \frac{2,163 \times 10 \times 8}{16} = 10,813\%$$

- Konsentrasi Etil Asetat Rimpang= 400 $\mu\text{g/mL}$, sampel = 16mg/6,66mL

1. 0,4171

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4171-0,1453}{0,1033} = \frac{0,2718}{0,1033} = 2,631 \text{ mg/mL}$$

2. 0,4174

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4174-0,1453}{0,1033} = \frac{0,2721}{0,1033} = 2,634 \text{ mg/mL}$$

3. 0,4176

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4176-0,1453}{0,1033} = \frac{0,2723}{0,1033} = 2,636 \text{ mg/mL}$$

- % Kadar Etil Asetat Rimpang

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,631 \times 10 \times 6,66}{16} = 10,952\%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,634 \times 10 \times 6,66}{16} = 10,964\%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{2,636 \times 10 \times 6,66}{16} = 10,972\%$$

- Konsentrasi Etanol Rimpang = 600µg/mL, sampel = 24mg/10mL
 1. 0,4808

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4808-0,1453}{0,1033} = \frac{0,3355}{0,1033} = 3,248 \text{ mg/mL}$$
 2. 0,4815

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4815-0,1453}{0,1033} = \frac{0,3362}{0,1033} = 3,255 \text{ mg/mL}$$
 3. 0,4848

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,4848-0,1453}{0,1033} = \frac{0,3395}{0,1033} = 3,287 \text{ mg/mL}$$
- % Kadar Etanol Rimpang
 1. % Kadar = $\frac{C \times V \times FP}{M} = \frac{3,248 \times 10 \times 10}{24} = 13,533\%$
 2. % Kadar = $\frac{C \times V \times FP}{M} = \frac{3,255 \times 10 \times 10}{24} = 13,561\%$
 3. % Kadar = $\frac{C \times V \times FP}{M} = \frac{3,287 \times 10 \times 10}{24} = 13,694\%$
- Konsentrasi N-heksan Daun = 400 µg/mL, sampel = 16mg/10mL
 1. 0,2636

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2636-0,1453}{0,1033} = \frac{0,1183}{0,1033} = 1,145 \text{ mg/mL}$$
 2. 0,2805

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2805-0,1453}{0,1033} = \frac{0,1352}{0,1033} = 1,309 \text{ mg/mL}$$
 3. 0,2823

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,2823-0,1453}{0,1033} = \frac{0,1370}{0,1033} = 1,326 \text{ mg/mL}$$
- % Kadar N-heksan Daun
 1. % Kadar = $\frac{C \times V \times FP}{M} = \frac{1,145 \times 10 \times 10}{16} = 7,156 \%$
 2. % Kadar = $\frac{C \times V \times FP}{M} = \frac{1,309 \times 10 \times 10}{16} = 8,181 \%$
 3. % Kadar = $\frac{C \times V \times FP}{M} = \frac{1,326 \times 10 \times 10}{16} = 8,287 \%$
- Konsentrasi Etil Asetat Daun = 600 µg/mL, sampel = 24mg/8,6mL
 1. 0,5714

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5714-0,1453}{0,1033} = \frac{0,4261}{0,1033} = 4,125 \text{ mg/mL}$$
 2. 0,5733

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5733-0,1453}{0,1033} = \frac{0,4280}{0,1033} = 4,143 \text{ mg/mL}$$
 3. 0,5759

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5759-0,1453}{0,1033} = \frac{0,4306}{0,1033} = 4,168 \text{ mg/mL}$$
- % Kadar Etil Asetat Daun

$$\begin{aligned}
 1. \% \text{ Kadar} &= \frac{C \times V \times FP}{M} = \frac{4,125 \times 10 \times 8,6}{24} = 14,781 \% \\
 2. \% \text{ Kadar} &= \frac{C \times V \times FP}{M} = \frac{4,143 \times 10 \times 8,6}{24} = 14,847 \% \\
 3. \% \text{ Kadar} &= \frac{C \times V \times FP}{M} = \frac{4,168 \times 10 \times 8,6}{24} = 14,937 \%
 \end{aligned}$$

- Konsentrasi Etanol Daun = 400µg/mL, sampel = 16mg/5,7mL

1. 0,5592

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5592-0,1453}{0,1033} = \frac{0,4139}{0,1033} = 4,007 \text{ mg/mL}$$

2. 0,5594

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5594-0,1453}{0,1033} = \frac{0,4141}{0,1033} = 4,009 \text{ mg/mL}$$

3. 0,5599

$$\text{Konsentrasi } (x) = \frac{y+a}{b} = \frac{0,5599-0,1453}{0,1033} = \frac{0,4146}{0,1033} = 4,014 \text{ mg/mL}$$

- % Kadar N-heksan Daun

$$1. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,007 \times 10 \times 5,7}{16} = 14,299 \%$$

$$2. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,009 \times 10 \times 5,7}{16} = 14,306 \%$$

$$3. \% \text{ Kadar} = \frac{C \times V \times FP}{M} = \frac{4,014 \times 10 \times 5,7}{16} = 14,323\%$$

Lampiran 4. Kolerasi Fenol Dan Flavonoid Total Dengan Aktivitas Antibakteri

1. Ekstrak Rimpang Kunyit

Correlations

		Flavonoid Rimpang	Aktivitas Bakteri Rimpang
Flavonoid Rimpang	Pearson Correlation	1	-.346
	Sig (2-tailed)		.775
	N	3	3
Aktivitas Bakteri Rimpang	Pearson Correlation	-.346	1
	Sig (2-tailed)	.775	
	N	3	3

Correlations

		Flavonoid Daun	Aktivitas Bakteri Daun
Flavonoid Daun	Pearson Correlation	1	.972
	Sig (2-tailed)		.151
	N	3	3
Aktivitas Bakteri Daun	Pearson Correlation	.972	1
	Sig (2-tailed)	.151	
	N	3	3

2. Ekstrak Rimpang Daun

Correlations

		Fenol Rimpang	Aktivitas Bakteri Rimpang
Fenol Rimpang	Pearson Correlation	1	.599
	Sig (2-tailed)		.591
	N	3	3
Aktivitas Bakteri Rimpang	Pearson Correlation	.599	1
	Sig (2-tailed)	.591	
	N	3	3

Correlations

		Fenol Daun	Aktivitas Bakteri Daun
Fenol Daun	Pearson Correlation	1	-.298
	Sig (2-tailed)		.808
	N	3	3
Aktivitas Bakteri Daun	Pearson Correlation	-.298	1
	Sig (2-tailed)	.808	
	N	3	3

PROFIL PENULIS



ANASTASIA NOVIANI HANDOKO

Alamat: Dusun Kepuh RT 02 RW 01 Desa Legonkulon Kec. Legonkulon
Kota Subang Jawa Barat
kode pos : 41254
Nomor HP: 083148101156
Email: 201ff03063@bku.ac.id

PROFIL PRIBADI

Lulusan baru Farmasi. Memiliki jiwa leadership yang tinggi, siap bekerja dibawah tekanan, Memiliki pengalaman bergabung di beberapa organisasi dan bidang umum. Mencari peluang yang memberikan diri semakin berkembang.

KEAHLIAN

Public Speaking	Bekerjasama dalam Team
Penyuluh Kesehatan	Leadership

RIWAYAT PENDIDIKAN

SEKOLAH MENENGAH ATAS (SMA)

MIS DARUSSALAM KASOMALANG	2017 - 2020
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Perguruan Tinggi

Universitas Bhakti Kencana	2020 - 2024
Sarjana Farmasi	
Indeks Prestasi Kumulatif (IPK): 3,13	

RIWAYAT ORGANISASI

ANGGOTA PRAMUKA	2017 – 2019
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- BANTARA

PENGURUS OPPD BAGIAN PENERIMAAN TAMU	2018 – 2019
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- Sekretaris Penerimaan Tamu

PELATIHAN

BMS Global Softskill Development Training	Januari 2023
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